

PRACTICE 2 KEY – CHEM TEAM MOLE RATIOS

- 1) $\frac{3 \text{ mol H}_2}{1 \text{ mol N}_2}$ and $\frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3}$
- 2) $\frac{2 \text{ mol SO}_3}{1 \text{ mol O}_2}$ and $\frac{2 \text{ mol SO}_2}{1 \text{ mol O}_2}$
- 3) $\frac{1 \text{ mol Cl}_2}{1 \text{ mol PCl}_3}$ and $\frac{1 \text{ mol PCl}_5}{1 \text{ mol PCl}_3}$
- 4) $\frac{2 \text{ mol N}_2}{4 \text{ mol NH}_3}$ and $\frac{3 \text{ mol O}_2}{6 \text{ mol H}_2\text{O}}$
- 5) $\frac{3 \text{ mol CO}_2}{3 \text{ mol CO}}$ and $\frac{3 \text{ mol CO}}{2 \text{ mol Fe}}$
- 6) $5.00 \text{ mol O}_2 \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} = \boxed{10.0 \text{ mol H}_2\text{O}}$
- 7) $3.00 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}} = \boxed{1.50 \text{ mol O}_2}$
- 8) $3.00 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = \boxed{3.00 \text{ mol H}_2}$
- 9) $12.00 \text{ mol KClO}_3 \times \frac{3 \text{ mol O}_2}{2 \text{ mol KClO}_3} = \boxed{18.00 \text{ mol O}_2}$
- 10) $\frac{2 \text{ mol C}_4\text{H}_{10}}{13 \text{ mol O}_2}$
- 11) $\frac{13 \text{ mol O}_2}{8 \text{ mol CO}_2}$
- 12) $\frac{13 \text{ mol O}_2}{10 \text{ mol H}_2\text{O}}$
- 13) $\frac{2 \text{ mol C}_4\text{H}_{10}}{8 \text{ mol CO}_2}$
- 14) $\frac{2 \text{ mol C}_4\text{H}_{10}}{10 \text{ mol H}_2\text{O}}$